

Template-free synthesis of novel, highly-ordered 3D hierarchical Nb₃O₇(OH) superstructures with semiconductive and photoactive properties†

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3D hierarchical Nb₃O₇(OH) mesocrystals can be formed by self-organization from nanometer sized building blocks. The present study focuses on the synthesis and detailed investigation of mesocrystals, which can be achieved from a one-step, template-free hydrothermal synthesis approach. The obtained cubic superstructures consist of a periodic nanowire-network and combine a large surface area, high crystallinity, with a band gap of 3.2 eV and photocatalytic activity. Their easy processability in combination with the named excellent properties makes them promising candidates for a large number of applications. These include photochemical and photophysical devices where the Nb₃O₇(OH) mesocrystals can be used as electrode material since they are semiconducting and possess a large surface area. Generally the forces involved in the self-organized formation of mesocrystals are not fully understood. In this regard, the assembly of the Nb₃O₇(OH) mesocrystals was investigated in-depth applying transmission electron microscopy, scanning electron microscopy, UV/Vis measurements and electron energy-loss spectroscopy. Based on the achieved results a formation mechanisms is proposed, which expands the number of mechanisms for mesocrystal formation reported in literature. In addition, our study reveals different types of nanowire junctions and investigates their role at the stabilization of the networks.

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1. Introduction

Nanostructured materials find application in various fields, like functionalized nanoparticles in medicine¹ or porous electrodes in photovoltaics.² Next to the synthesis of nanoparticles and one-dimensional nanostructures of different materials much effort is made for the generation of more complex but highly-ordered three-dimensional superstructures. The developed synthesis strategies can generally be divided into template-based and non-template based methods. In connection with

this, Cölfen *et al.* introduce the term mesocrystal, which describes the self-assembly of non-spherical building blocks to ordered superstructures on the scale of several hundred nanometers to micrometers.^{3,4} To date, the self-organized formation of these mesocrystals is not fully understood. However, several models were proposed in literature, based on different driving forces like external and internal fields, organic matrixes or interparticle forces.⁵⁻⁷

Solvothermal synthesis methods, which enable the fabrication of superstructures, belong to the class of non-template based methods. They utilize high pressures for the crystallization of materials at comparably low temperatures. These reaction conditions apparently favor the self-organized formation of nanostructures. Reports show solvothermally achieved structures like CuInS₂-spheres,⁸ hierarchical Co₃O₄ superstructures⁹ or TiO₂-nanowires.¹⁰ TiO₂ is one of the most applied materials in the field of green energy and the decomposition of hazardous materials,^{11,14,15} some of the major tasks of this century. In general, materials applied in this field must fulfill certain preconditions like large accessible surface areas, long-range electronic connectivity, structural coherence and a suitable band structure.¹² Next to the most commonly applied TiO₂ and ZnO various other materials are investigated.¹³ Niobium oxide is promising as it exists in many polymorphic phases with different oxidation numbers and crystal structures, most

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featuring a high chemical stability, corrosion resistance in both acidic and alkaline media and, dependent on the crystal phase, band gaps in the range of 3.1 eV¹⁶ to 4.9 eV.¹⁷ Several studies utilize solvo-thermal methods to achieve nanostructured niobium oxide,^{18–21} whose application in dye-sensitized solar cells (DSSC)^{22–24} results in power conversion efficiencies up to 7%.²⁵ Besides being used in DSSCs, niobium oxide nanostructures find application in the field of electrochromics,^{26,27} catalysis,²⁸ field-emission displays,^{29,30} gas sensing³¹ and lithium batteries.³²

Here we report on novel three-dimensional hierarchical Nb₃O₇(OH) structures, which were synthesized *via* a template-free, one-step hydrothermal approach. The focus of our study lies on the formation of these hierarchical structures by analyzing intermediate growth states combining X-ray diffraction (XRD), scanning electron microscopy (SEM) and UV/Vis spectroscopy. This provides deeper insights into the self-organized formation of mesocrystals. Further analyses were performed at the atomic scale to get a better understanding of the arrangement of these hierarchical structures. For this purpose transmission electron microscopy (TEM), electron energy-loss spectroscopy (EELS) and scanning transmission electron microscopy (STEM) were performed. In addition to the structural analysis the presented structures prove to be viable for photochemical applications as shown in bleaching experiments with different dye molecules.

2. Results and discussion

2.1 Hydrothermal conditions enable the self-organized formation of Nb₃O₇(OH) superstructures

The applied hydrothermal synthesis strategy yields different well-controllable cube morphologies without using a template. The one-step synthesis approach is based on the hydrothermal treatment of an aqueous Nb(IV)chloride solution at temperatures between 150 and 200 °C. In dependence of the applied synthesis parameters (pH-value, temperature, synthesis time) a variety of cube morphologies can be achieved. These morphologies reach from compact to hollow cubes, with differently shaped surfaces. Amongst those, hollow cubes with surfaces consisting of nanowires stand out, as this unique morphology combines a large surface area (BET 53 m² g⁻¹) and a high crystallinity (Fig. 1). Consequently, the investigations focus on these superstructures aiming at a better understanding of their formation, stabilization and applicability. These nanowire-based hollow cubes can be achieved purely using a pH of -0.8, a synthesis temperature of 200 °C and a growth time of 24 hours.

2.2 Analysis of the mesocrystal formation

For a better understanding of mechanisms involved in the formation of these novel superstructures, the intermediate steps of the cubes formation process were analyzed. The SEM images displayed in Fig. 1 show the morphology of the cubes depending on the reaction time. Note that cubes form in the reaction solution already at the initial growth state at room temperature. These cubes are characterized by blunt edges,

which sharpen during the first hours of the hydrothermal synthesis. Substantial changes of the cube morphology were observed for a synthesis time of eight hours. The corresponding sample contains two morphologies emblematic for two different growth states: on the one hand tiny nanowires, which seem to lie loosely on the surface of the cubes and on the other hand larger nanowires occur. For synthesis times longer than eight hours the cubes are completely built up from nanowire networks, which grow at the expense of the material in the interspaces between the nanowires. The continuous changing of the morphology is accompanied by increasing mean cube sizes, reaching their maximum of 3.9 ± 1.1 μm for a synthesis time of 24 hours. For very large synthesis times (72 hours) the stability of the cubes decreases, resulting in wreckage of nanowires and a decreased mean cube size of 2.9 ± 0.6 μm, which indicates a preferential breakdown of larger cubes (quantitative analysis see ESI Fig. S1†).

Changes of the crystallographic properties which go along with the changing morphologies were analyzed using XRD (Fig. 1). The analysis shows that the samples are X-ray amorphous for shorter synthesis times, as the corresponding patterns only contain broad features attributable to short-range order phenomena. Crystallization of the sample starts after a synthesis time of two hours. For longer synthesis duration the observed reflexes become more pronounced. Rietveld analyses performed on the crystalline samples indicate that they all consist of pure Nb₃O₇(OH) phase. The changing crystallographic properties are also reflected in the band gap of the material, which was determined from UV/Vis data based Tauc plots. Investigation of the band gap of the different morphologies shows a slightly decreasing trend for increasing synthesis time, spanning an energy range of 0.14 eV. Only the band gap of the cubes obtained from a synthesis time of one hour breaks from this trend, as it is 0.2 eV larger than the band gap of the amorphous starting phase (0 h sample) (see ESI Fig. S2 and Table S1†). The morphologic changes during the hydrothermal treatment coincide with a chemical reaction resulting in the formation of niobium oxide out of niobium chloride. The molecular formula of Nb₃O₇(OH) indicates the presence of niobium(V) in the cubes. Analysis of the bonding characteristic by EELS measurements in the TEM was performed to verify this oxidation state. Comparing the O-K-edge and the plasmon of EEL spectra detected for single nanowires with spectra from literature³³ indicates the pentavalence of niobium in Nb₃O₇(OH), as the shape of the accordant edges agree with the ones reported for Nb₂O₅ (see ESI Fig. S3†). This reveals the oxidation reaction happening with the formation process of the cubes, resulting in the transformation of Nb(IV)chloride to Nb(V)oxide as given in the following equation:



Combining the results of these different analyses allows to set up a model of the formation mechanism of these cubes. A schematic drawing illustrating the growth is shown in Fig. 1. The growth starts from amorphous, hollow cubes with smooth



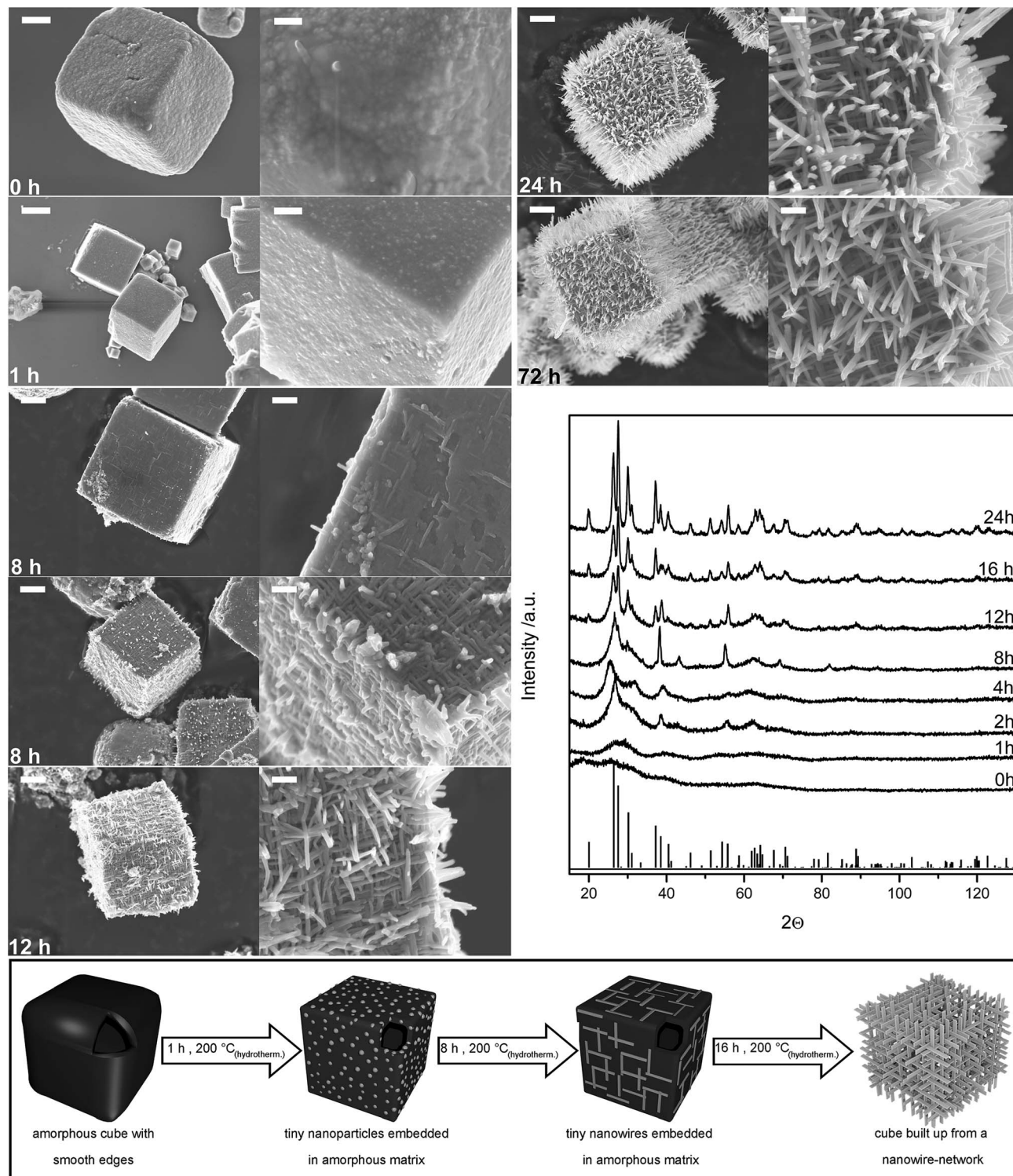


Fig. 1 SEM images of Nb₃O₇(OH) cube morphologies obtained for different synthesis times. The left row shows an overview SEM image of single cubes for different times of hydrothermal treatment – 0 h, 1 h, 8 h, 12 h, 24 h, 72 h – (scale bar 1 μm) and the right row shows a zoom in on one edge of the corresponding cube (scale bar 200 nm). The development of nanowires during the hydrothermal treatment can be observed. XRD pattern of the different morphologies and the reference data of Nb₃O₇F (JCPDS, card no. 74-2363), which is isomorphic to Nb₃O₇(OH),³⁷ can be seen in the graph on the right. At the bottom a schematic drawing is shown, which summarizes the intermediate steps of the hydrothermal growth of the Nb₃O₇(OH) superstructures.

surfaces, which already form at room temperature. With beginning of the hydrothermal treatment apparently tiny nanoparticles form, which are embedded in an amorphous

matrix. These nanoparticles cause a quantum confinement effect resulting in an increased band gap. Comparable results were reported by Anpo *et al.*, who observed a blue shift of



0.16 eV for the band gap of anatase TiO₂ nanoparticles of 38 Å diameter compared to bulk anatase due to quantum confinement effects.³⁴ The Nb_xO_y particles most likely serve as seed crystals for the formation of the nanowires *via* one-dimensional crystal growth. The formation of nanowires due to one-dimensional crystal growth was also reported to occur under hydrothermal conditions for rutile TiO₂.¹⁰ Density functional theory calculations performed for rutile show that one-dimensional growth is thermodynamically favored as the size of high-energy crystal surfaces is reduced for the benefit of lower-energy surfaces.^{35,36} Similar reasons might cause the one-dimensional growth of Nb₃O₇(OH). Here, the one-dimensional growth starts with tiny nanowires, which are embedded in an amorphous matrix (observed for synthesis times of eight hours). With ongoing growth (for synthesis times larger than 16 hours), the nanowires increase in size at the expense of the amorphous material in the interspaces. These findings suggest the involvement of an Ostwald ripening process during the wire formation. Ostwald ripening was also reported to be the reason for the formation of hollow Nb₂O₅ spheres by Xue *et al.*²¹ and Li *et al.*³⁸ Some steps of the mechanism found for the formation of the Nb₃O₇(OH) mesocrystals are in compliance with models of the self-organized formation of mesocrystals reported in literature.^{4,39} Similarity is found to a four-step mechanism proposed by Gong *et al.*⁴ and the formation mechanism reported by Zhang *et al.*⁴⁰ The mechanism of Gong *et al.* is subdivided into the initial nucleation, the self-assembly *via* oriented aggregation, crystal growth *via* Ostwald ripening and a subsequent separation process.⁴ Zhang *et al.* show the formation of nanostructured spheres starting from nanoparticles which transform to spheres built up of nanoparticles. With ongoing hydrothermal treatment first interlaced connected streak structures form, which are favored for further growth of nanosheets.⁴⁰

2.3 Investigation of the structure of the mesocrystals at atomic scale

For a better understanding of their structure, the morphology of the cubes consisting of a nanowire arrangement was investigated intensively. SEM (Fig. 2a–c) and TEM (Fig. 2d) images of the cubes show that their walls are composed of nanowire networks. TEM images of network fragments reveal a nearly perfect perpendicular arrangement of the nanowires in the network (Fig. 2e). This regularity is reflected in the electron diffraction pattern of the network (Fig. 2f), which consists of distinct reflexes usually characteristic for single crystalline materials. However, TEM images reveal black contours within the wires, attributable to the presence of strain and defects in the network (Fig. 2e). SEM images of broken cubes confirm the hollowness of the cubes. Cross-sections of the cube walls exhibit additional nanowires oriented perpendicular to the cube edges (Fig. 2b and c). These nanowires appear as dark spots in TEM images of the network, because they are aligned parallel to the beam direction (three are marked with black arrows in Fig. 2e).

Energy-dispersive X-ray (EDX) analyses performed in the TEM on single nanowires, which broke off from cubes

synthesized for 24 hours (Fig. 3a), show that the nanowires homogeneously consist of niobium and oxygen (EDX data see ESI Fig. S4†). Electron diffraction experiments performed for the respective nanowire indicate that it is single crystalline Nb₃O₇(OH) phase and grows in the ⟨010⟩ direction (the indexing was verified with simulations using the JEMS software,⁴¹ see ESI Fig. S5a and b†). Comparable results were reported by Zhang *et al.*, who investigated Nb₃O₇(OH) nanowire arrays, which did not form three-dimensional assemblies.²⁵ Annular dark field images (ADF) detected for one single nanowire reflect the crystal symmetry, which is characterized by [NbO₆][−] octahedra connected tip- and edge-wise. A schematic drawing of the crystal structure showing these octahedra is included in the ADF image in Fig. 3b. Furthermore, it is worth noticing that the ⟨001⟩ zone axis is always observed for single nanowires lying on a holey carbon film, indicating an enlarged ⟨001⟩ surface compared to the ⟨100⟩ surface (see schematic drawing of a nanowire in Fig. 3c).

Analysis of the nanowire network reveals that it is built up from nanowires, all facing the same crystallographic orientation, displaying the ⟨100⟩ zone axis (the indexing of one single nanowire in the network was verified by simulations, see ESI Fig. S5c and d†). In this regard, the diffraction pattern of the network shown in Fig. 2f was indexed as the superposition of the diffraction pattern of two nanowires rotated by 90° with respect to each other (detailed analysis see ESI Fig. S6†). Fast Fourier transformation (FFT) analyses of HRTEM images confirm this indexing (see ESI Fig. S7a and b†). In addition, these HRTEM images of the network exhibit two different types of nanowire junctions, which were further analyzed using site selective EELS based thickness measurements and scanning transmission electron microscopy (STEM). One junction type results from a T-shaped arrangement of two nanowires. As no thickness changes could be detected at these junctions, the corresponding nanowires do not seem to overlap (see Fig. 3d). The second junction type results from two nanowires overlaying each other, forming a crossing. The enhanced thickness at the overlap region, which corresponds to about 80% of the thickness sum of the respective nanowires, shows that the nanowires partly interpenetrate each other at the overlap region (see Fig. 3e and ESI Fig. S7c–f†). FFT analyses reveal that these overlap regions cause additional reflexes appearing in the diffraction pattern of the network (marked with white circles in Fig. 2f). The intensity of those reflexes depends on the network size, which corresponds to the quantity of crossings, in that way that it increases with growing network size.

The analysis indicates a faceted shape of the nanowires with preferred {001} planes. These nanowires form the building blocks of 3D hierarchical cubes, which consist of nanowire-networks. The high degree of regularity of the nanowire arrangement results in a single crystalline appearance of the complex superstructures. This, and the fact that they form self-organized from solution proposes them categorizable as mesocrystals following the definition of Cölfen *et al.*³ Neither the formation nor the stabilization of mesocrystals is generally understood in literature. The aforementioned results indicate a major role of the nanowire crossings at the stabilization of the



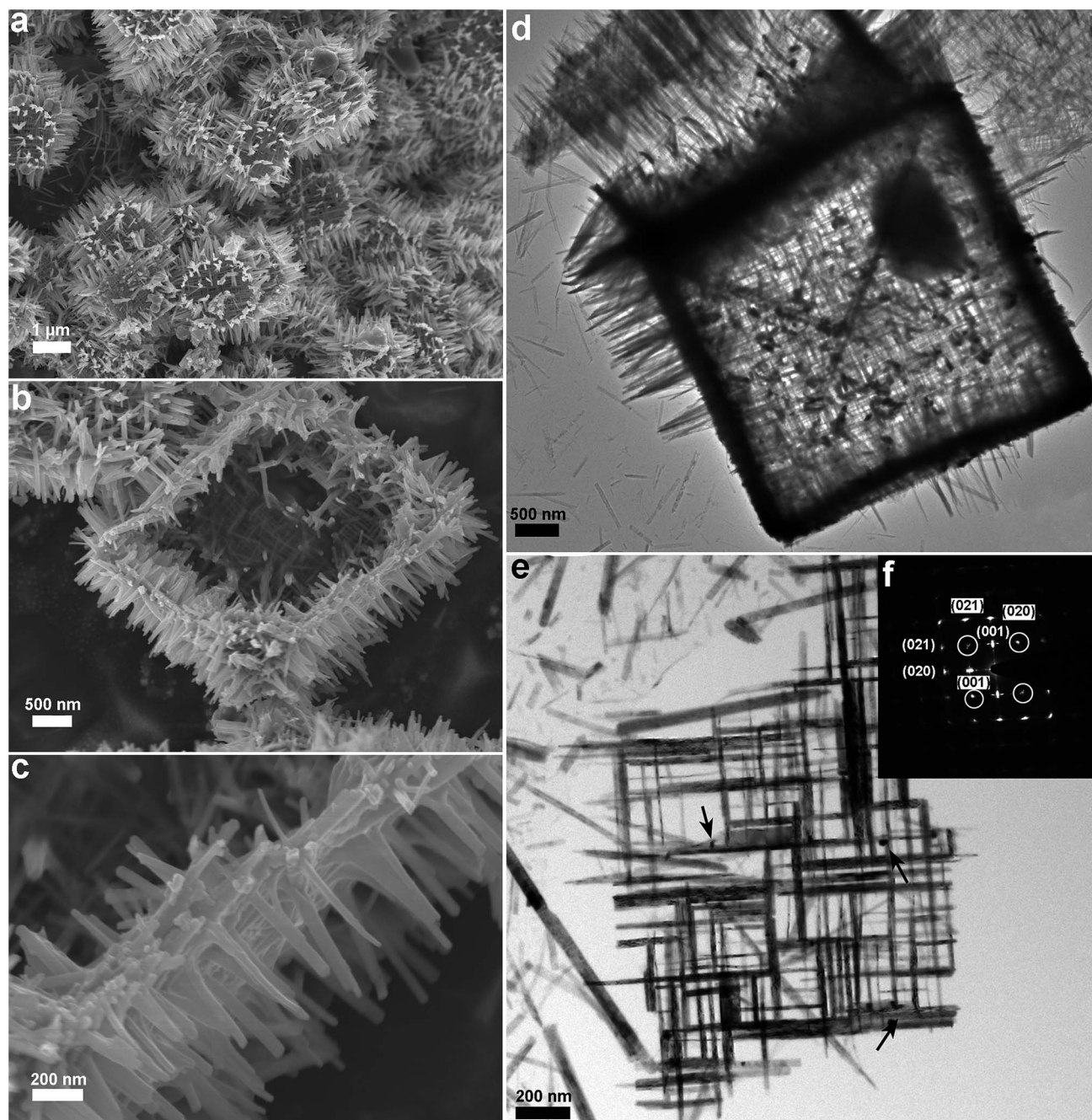


Fig. 2 Electron microscopic investigation of the $\text{Nb}_3\text{O}_7(\text{OH})$ mesocrystals. (a) An overview SEM image of the sample. (b) SEM image of one cube with its top broken off, a higher magnification of the cube wall can be seen in (c). (d) Overlap TEM image of a topless cube lying on a carbon grid. (e) TEM image of a fragment of one cube wall and the corresponding electron diffraction pattern (f), which is indexed as combined pattern of two nanowires.

networks, as a partial interpenetration of the nanowires could be shown for these junctions *via* site-selective thickness measurements. The resulting increased thickness at the overlap region compared to plain nanowires is also reflected in the diffraction patterns of the network, which show additional reflexes most likely attributable to double diffraction. However, the partial interpenetration of the nanowires at the crossing is apparently enforced by the orthorhombic crystal symmetry of $\text{Nb}_3\text{O}_7(\text{OH})$ which is characterized by $a = 20.740 \text{ \AA}$, $b = 3.823 \text{ \AA}$

and $c = 3.936 \text{ \AA}$.³⁷ The small lattice mismatch of the b and c direction enables the preservation of the atomic columns in $\langle 100 \rangle$ direction at the overlap regions of the crossing, thus the octahedral coordination of niobium can be preserved (see ESI Fig. S7f†). Nonetheless, the small difference between the two crystal directions causes strain in the network, which is visible in bright-field images (Fig. 2e). Yet, the driving forces involved in the perpendicular alignment of the nanowires are still unknown. Different polarities of the crystal surfaces might hint



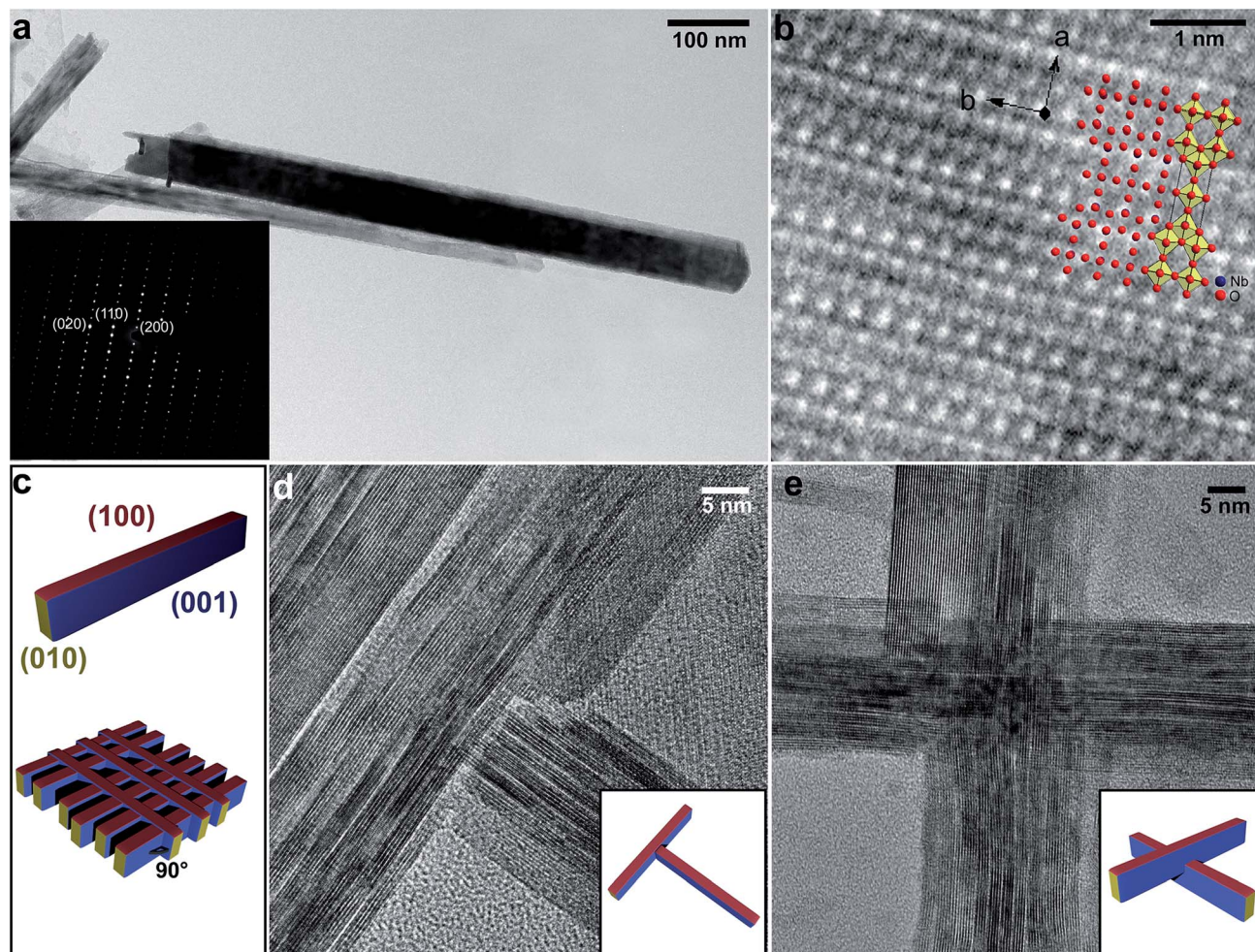


Fig. 3 Electron microscopic investigation of one single $\text{Nb}_3\text{O}_7(\text{OH})$ nanowire on a holey carbon film, which broke off from a network, and of the different nanowire junctions in the network. (a) Overlap bright field image of the investigated nanowire and corresponding electron diffraction pattern detected for this wire, which is indexed for $\text{Nb}_3\text{O}_7(\text{OH})$ in $\langle 001 \rangle$ zone axis (inset). (b) ADF image of a single nanowire including a schematic drawing of the crystal symmetry. (c) Schematic drawing illustrating the crystal shape of the nanowires and the crystallographic arrangement of the nanowires in the network. (d) HRTEM image of a T-shaped nanowire junction and schematic drawing showing the arrangement of the nanowires at the junction (inset). (e) HRTEM image of a nanowire crossing and schematic drawing of the junction (inset).

to an explanation,⁵ as buck charging of the (010) and (100) faces would explain the perpendicular arrangement due to balanced repulsion and attraction. The mechanism of formation of the additional nanowires, standing like spikes from the edges of the cubes, and their role in the stabilization of the mesocrystals is not understood yet and will be the subject of future investigations.

2.4 Adequacy of the $\text{Nb}_3\text{O}_7(\text{OH})$ mesocrystals as new electrode material

The mesocrystals benefit from their large surface area, high crystallinity and direct electron transport paths. In this regard their optoelectronic properties related to their application in photochemical devices were investigated. The band gap of a material is one important factor. Thus, UV/Vis data based Tauc plots (Fig. 4a) were applied. They show a band gap of 3.19 ± 0.05 eV for the sample synthesized for 24 hours. Since

UV/Vis spectroscopy averages over the whole sample, additional site-selective EELS measurements were performed in the TEM to determine the band gap of one single nanowire of the same sample. The corresponding low-loss EEL spectrum is shown in Fig. 4b after subtraction of the zero-loss peak. It confirms the size of the band gap, yielding a value of 3.2 ± 0.1 eV. However, the determined band gaps are slightly larger than the one reported by Wu *et al.*, who found a band gap of 3.03 eV for their $\text{Nb}_3\text{O}_7(\text{OH})$ thin films.¹⁶ The 2D photoluminescence spectrum of the cubes grown for 72 hours (Fig. 4c) shows one major feature with an excitation maximum in the range of 4.1 to 3.1 eV and an emission range of 2.8 to 2.5 eV. This feature can be attributed to interband transitions from the valence to the conduction band, as the energy corresponds to the size of the band gap and electronic excitations to higher energetic states, thus likewise confirming the determined band gap. However, the emission range points towards a possible existence of trap states.



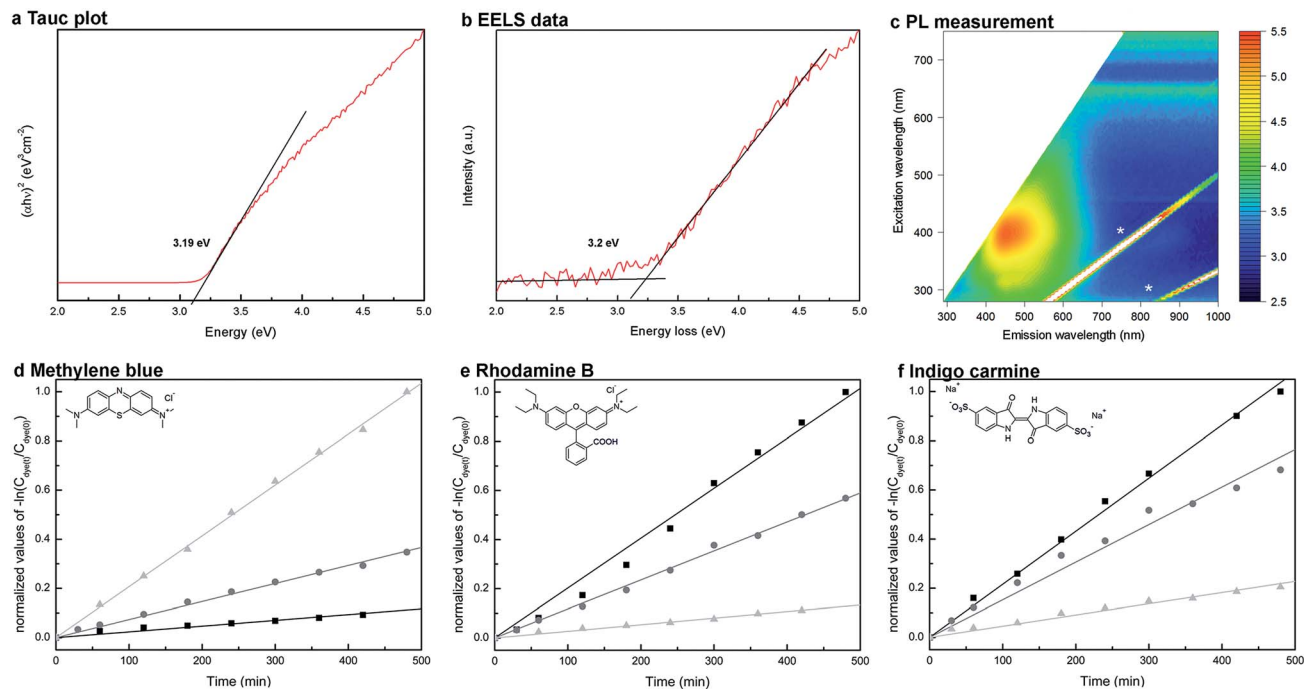


Fig. 4 Measurements of the optoelectronic properties of the Nb₃O₇(OH) mesocrystals. (a) UV/Vis data based Tauc plot, from which a band gap of 3.19 ± 0.05 eV was determined for cubes achieved for a synthesis time of 24 hours. (b) Low loss EEL spectrum detected at a single nanowire of the same sample (24 hours) after removing the zero-loss peak. The linear fit method⁴² was used to determine the band gap. The spectrum yields a band gap of 3.2 ± 0.1 eV. (c) 2D photoluminescence spectrum of a sample synthesized for 72 hours, showing one intensive absorption peak between 400 and 500 nm. Emission intensity is shown in arbitrary units on a logarithmic scale. Spectral features labeled with * are experimental artifacts at the doubled and tripled excitation wavelengths and do not originate from the sample. (d–f) Measurement of the photocatalytic degradation of three different dyes at three different pH values (pH 2 (■), pH 6 (●) and pH 10 (▲)). The kinetic rate constant can be determined from the graph obtained by plotting $-\ln(C_{\text{dye}}/C_{\text{dye}0})$ versus the irradiation time t . The corresponding graphs are shown in (d) for methylene blue, in (e) for rhodamine B and in (f) for indigo carmine.

The band gap of Nb₃O₇(OH) is smaller or in the range of common other metal oxides, like TiO₂, Nb₂O₅ or ZnO which are used as electrode material in photovoltaics and photocatalysis.¹³ The photocatalytic degradation of dyes is an elegant method to investigate their suitability for photochemical applications. In the present study the analysis of three dyes with different colors and functionality shows a successful degradation. However, a strong pH-value dependence of the kinetic degradation rate was observed (the kinetic rates correspond to the slope of the graphs summarized in Fig. 4d–f). While the adsorption of indigo carmine and methylene blue occurs for specific pH values, this was not the case for rhodamine B. For indigo carmine adsorption was best at a pH of 2, which is reflected in the highest photocatalytic degradation rate at this pH value, while methylene blue was best adsorbed at a pH of 10 (see ESI Fig. S8d†). Nevertheless, the investigations reveal that the adsorption of the dye is not necessary for the photocatalytic degradation, but increases its reaction kinetic. The photocatalytic degradation of rhodamine B by Nb₃O₇(OH) films was also shown by Wu *et al.*¹⁶ The pH dependence of its photocatalytic degradation might be related to the pH dependent intermediate formation of either hydroxyl or hydrogen radicals.⁴³ These findings, namely an attractive band gap and photocatalytic activity, indicate the suitability of these Nb₃O₇(OH) mesocrystals as electrode material in functional devices.

Correspondent testing goes beyond the scope of this work, but Zhang *et al.* already showed the successful integration of Nb₃O₇(OH) nanowire arrays into DSSC reaching efficiencies up to 7%.²⁵

3. Conclusion

This study reports on novel Nb₃O₇(OH) mesocrystals with hollow nanostructured cubic morphologies, which are grown in a one-step, template-free hydrothermal synthesis. Their self-organized assembly was investigated showing that initially amorphous hollow cubes with smooth surfaces form, which transform to cubes built up from crystalline Nb₃O₇(OH) nanowire-networks during the hydrothermal synthesis. The morphologic changes go along with an oxidation reaction resulting in the formation of niobium(v)oxide. From our results we motivate a model for a new type of crystal growth which has not been observed before in literature and expands the numbers of formation mechanisms reported for other mesocrystals.

The construction of the walls was investigated at atomic scale, showing that the walls of the hollow cubes consist of a regular nanowire network. In the network the nanowires are aligned in a perpendicular fashion. This regularity is also reflected in the electron diffraction pattern, which shows distinct reflexes. The self-organized formation and the high



degree of regularity indicate that these superstructures fulfill the criteria for mesocrystals following the definition given by Cölfen *et al.*³ Their stabilization is most likely related to the interpenetration of the nanowires at crossings, which at the same time triggers the perpendicular arrangement.

In addition the properties of the cubes, namely their large surface area, their band gap of 3.2 eV, their high chemical stability and photochemical activity render them attractive for photochemical and photophysical devices. Furthermore the reported synthesis strategy and mechanistic investigations support the development of mesocrystals of various materials and the understanding of the forces triggering self-assembly.

4. Experimental section

4.1 Synthetic procedures

All chemicals were used as supplied without further purification. For the synthesis of the cubes niobium(IV)chloride–tetrahydrofuran complex (0.2 g, 0.5 mmol; supplied by Sigma Aldrich) was mixed with concentrated HCl (5 mL; 37% HCl, p.a., Brenntag). To this solution 5 mL H₂O were added under vigorous stirring. The solution was transferred into a Teflon liner in a stainless steel autoclave and placed into an oven at 200 °C. After cooling down to room temperature, the reaction product was collected and washed several times with ethanol to remove remaining starting material.

4.2 Materials and methods

Investigation of the morphology. Scanning electron microscopy was applied for the analysis of the morphology of the cubes in dependence of the synthesis time. A Zeiss Gemini Ultra Field Emission electron microscope equipped with an in-lens detector was used at 4 kV. For the analysis the powders were fixed on silicon wafers. To further reduce charging, some of the samples were sputtered with carbon using a BAL-TEC MED020 coating system.

UV/Vis measurements. The UV/Vis measurements were performed using an Agilent Cary 5000 spectrometer, detecting the absorption spectra in a wavelength range of 200 to 750 nm, using a step size of 1 nm. The powder samples were fixed on a quartz substrate. The detected UV/Vis data form the basis of the Tauc plots for direct semiconductors, which were used for the band gap determination.

TEM analysis. The samples were prepared from a sonicated mixture of the powder sample with pure ethanol. This mixture was dropped onto a copper grid with a holey carbon film. To reduce contamination the samples were plasma treated prior to the measurements using a Gatan Solarus Advanced Plasma (H₂/O₂ Model 950) System for 10 seconds. A FEI Titan 80-300 equipped with a Gatan Tridiem EELS detector and an EDAX ultra thin window Si(Li) EDX detector was used at 300 kV for the HRTEM, STEM and thickness determination *via* EELS. For the thickness determination EELS measurements were performed in STEM mode averaging over small sample areas (in the range of 6 × 6 nm²), using a collection semi angle of 15.6 mrad, a convergence semi angle of 9.5 mrad, a total acquisition time of

5 s and a dispersion of 0.1 eV per channel. The thickness was determined using the log-ratio-method implemented in the Digital micrograph software (Gatan).

Low loss EELS acquisition and Cs corrected STEM images were taken with a FEI Titan³ G2 60-300 equipped with an aberration-corrector for the probe-forming lens system, a high brightness field emission (XFEGTM) source, a monochromator and a high-resolution electron-energy loss spectrometer (Gatan 966 GIF) operated at 80 kV. Data were acquired in STEM spot mode and by scanning an area of about 20 × 20 nm². A dispersion of 0.025 eV per channel and total acquisition times up to 10 s were used. The convergence semi angle was 18.7 mrad, the collection semi angle 6.5 mrad. For the band gap measurement the zero-loss tail was removed by a power law fit. The band gap was determined using the linear fit method,⁴² estimating the band gap from the intersection of a straight line originating from the background level with a linear fit to the onset of the spectrum.

Photoluminescence measurements. Photoluminescence measurements were performed on a Jobin Yvon Fluorolog Spectrometer (Horiba) equipped with a 500 W Xenon lamp. The two dimensional spectrum was detected for an emission range of 250 to 1000 nm and excitation range of 240 to 750 nm using a step size of 5 nm and an illumination angle of 60°. For the measurements the sample was fixed onto a quartz substrate.

XRD analysis. XRD analysis was performed using a Seifert THETA/THETA-diffractometer (GE Inspection Technologies) equipped with a Meteor OD detector, using Co-K α radiation. A 2 θ detection range of 15–130°, a step size of 2 θ = 0.05° and a count time of 30 s per (2 θ -step) were used. The patterns were analyzed by Rietveld analysis using the Materials Analysis Using Diffraction (MAUD) software by Luca Lutterotti,⁴⁴ after removing the background.

Photocatalytic degradation of dyes. The degradation of three different dyes (rhodamine B, methylene blue, indigo carmine) was investigated at three different pH values (2, 6 and 10). For the measurements Nb₃O₇(OH) powder (0.5 mg, 0.001 mmol) was mixed with 3 mL aqueous dye solution (concentration: indigo carmine (57 mg L⁻¹), methylene blue (800 mg L⁻¹), rhodamine B (11 mg L⁻¹)). The absorption of the different dye solutions was measured using an Ocean Optics USB4000 spectrometer. To ensure the establishment of an adsorption-desorption equilibrium between Nb₃O₇(OH) and the dye, the mixtures were incubated for 1.5 hours in the dark. The catalytic degradation was started *via* UV irradiation using a UVP UVGL-58 handheld UV lamp at long-wave UV irradiation. The kinetic constant was determined assuming a pseudo-first order kinetic model as the slope of the graph achieved by plotting $-\ln(C_{\text{dye}}/C_0)$ versus the irradiation time *t*.

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Notes and references

- 1 R. Ladj, A. Bitar, M. Eissa, Y. Mugnier, R. Le Dantec, H. Fessi and A. Elaissari, *J. Mater. Chem. B*, 2013, **1**, 1381.
- 2 H. N. Kim and J. H. Moon, *ACS Appl. Mater. Interfaces*, 2012, **4**, 5821.
- 3 H. Cölfen and M. Antonietti, *Angew. Chem.*, 2005, **44**, 5576.
- 4 Q. Gong, X. Qian, X. Ma and Z. Zhu, *Cryst. Growth Des.*, 2006, **6**, 1821.
- 5 R. Q. Song and H. Cölfen, *Adv. Mater.*, 2010, **22**, 1301.
- 6 K. J. Bishop, C. E. Wilmer, S. Soh and B. A. Grzybowski, *Small*, 2009, **5**, 1600.
- 7 S. J. Sedlmaier, T. Dennenwaldt, C. Scheu and W. Schnick, *J. Mater. Chem.*, 2012, **22**, 15511.
- 8 A. S. Wochnik, A. Frank, C. Heinzl, J. Häusler, J. Schneider, R. Hoffmann, S. Matich and C. Scheu, *Solid State Sci.*, 2013, **26**, 23.
- 9 L. Zhang, X. Zhao, W. Ma, M. Wu, N. Qian and W. Lu, *CrystEngComm*, 2013, **15**, 1389.
- 10 B. Liu and E. S. Aydil, *J. Am. Chem. Soc.*, 2009, **131**, 3985.
- 11 X. Chen and S. S. Mao, *Chem. Rev.*, 2007, **107**, 2891.
- 12 P. Docampo, S. Guldin, U. Steiner and H. J. Snaith, *J. Phys. Chem. Lett.*, 2013, **4**, 698.
- 13 R. Jose, V. Thavasi and S. Ramakrishna, *J. Am. Ceram. Soc.*, 2009, **92**, 289.
- 14 M. D. Hernández-Alonso, F. Fresno, S. Suárez and J. M. Coronado, *Energy Environ. Sci.*, 2009, **2**, 1231.
- 15 M. Wang, J. Icozzia, L. Sun, C. Lin and Z. Lin, *Energy Environ. Sci.*, 2014, **7**, 2182.
- 16 J. Wu, J. Wang, H. Li and D. Xue, *Thin Solid Films*, 2013, **544**, 545.
- 17 G. Agarwal and G. B. Reddy, *J. Mater. Sci.: Mater. Electron.*, 2005, **16**, 21.
- 18 M. Wei, Z.-m. Qi, M. Ichihara and H. Zhou, *Acta Mater.*, 2008, **56**, 2488.
- 19 J. Liu, D. Xue and K. Li, *Nanoscale Res. Lett.*, 2011, **6**, 138.
- 20 C. Yan, L. Nikolova, A. Dadvand, C. Harnagea, A. Sarkissian, D. F. Perepichka, D. Xue and F. Rosei, *Adv. Mater.*, 2010, **22**, 1741.
- 21 D. Xue and F. Liu, *Mod. Phys. Lett. B*, 2009, **23**, 3769.
- 22 K. Sayama, H. Sugihara and H. Arakawa, *Chem. Mater.*, 1998, **10**, 7.
- 23 P. Guo and M. A. Aegerter, *Thin Solid Films*, 1999, **351**, 290.
- 24 R. Ghosh, M. K. Brennaman, T. Uher, M. R. Ok, E. T. Samulski, L. E. McNeil, T. J. Meyer and R. Lopez, *ACS Appl. Mater. Interfaces*, 2011, **3**, 3929.
- 25 H. Zhang, Y. Wang, D. Yang, Y. Li, H. Liu, P. Liu, B. J. Wood and H. Zhao, *Adv. Mater.*, 2012, **24**, 1598.
- 26 T. Maruyama and S. Arai, *Appl. Phys. Lett.*, 1993, **63**, 869.
- 27 N. Özer, D.-G. Chen and C. M. Lampert, *Thin Solid Films*, 1996, **277**, 162.
- 28 Y. Zhao, X. Zhou, L. Ye and S. C. E. Tsang, *Nano Rev.*, 2012, **3**, 1.
- 29 M. Žumer, V. Nemanič, B. Zajec, M. Remškar, A. Mrzel and D. Mihailović, *Appl. Phys. Lett.*, 2004, **84**, 3615.
- 30 B. Varghese, S. C. Haur and C.-T. Lim, *J. Phys. Chem. C*, 2008, **112**, 10008.
- 31 D. Rosenfeld, P. E. Schmid, S. Széles, F. Lévy, V. Demarne and A. Grisel, *Sens. Actuators, B*, 1996, **37**, 83.
- 32 A. L. Viet, M. V. Reddy, R. Jose, B. V. R. Chowdari and S. Ramakrishna, *J. Phys. Chem. C*, 2010, **114**, 664.
- 33 D. Bach, R. Schneider, D. Gerthsen, J. Verbeeck and W. Sigle, *Microsc. Microanal.*, 2009, **15**, 505.
- 34 M. Anpo, T. Shima, S. Kodama and Y. Kubokawa, *J. Phys. Chem.*, 1987, **91**, 4305.
- 35 H. Perron, C. Domain, J. Roques, R. Drot, E. Simoni and H. Catalette, *Theor. Chem. Acc.*, 2007, **117**, 565.
- 36 M. Ramamoorthy, D. Vanderbilt and R. King-Smith, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **49**, 16721.
- 37 F. Izumi and H. Kodama, *Z. Anorg. Allg. Chem.*, 1978, **441**, 8.
- 38 L. Li, J. Deng, R. Yu, J. Chen, Z. Wang and X. Xing, *J. Mater. Chem. A*, 2013, **1**, 11894.
- 39 T. Wang, H. Cölfen and M. Antonietti, *J. Am. Chem. Soc.*, 2005, **127**, 3246–3247.
- 40 H. Zhang, Y. Li, Y. Wang, P. Liu, H. Yang, X. Yao, T. An, B. J. Wood and H. Zhao, *J. Mater. Chem. A*, 2013, **1**, 6563.
- 41 P. A. Stadelmann, *JEMS-EMS java version*, École Polytechnique Fédérale de Lausanne, Lausanne, Switzerland, 2004.
- 42 J. Park, S. Heo, J. G. Chung, H. Kim, H. Lee, K. Kim and G. S. Park, *Ultramicroscopy*, 2009, **109**, 1183.
- 43 U. G. Akpan and B. H. Hameed, *J. Hazard. Mater.*, 2009, **170**, 520.
- 44 L. Lutterotti, *Nucl. Instrum. Methods Phys. Res., Sect. B*, 2010, **268**, 334.

